

AD-A240134

UNCLASSIFIED

SECURITY CLASSIFICATION OF THIS PAGE

REPORT DOCUMENTATION PAGE				Form Approved OMB No 0704-0188	
1a REPORT SECURITY CLASSIFICATION UNCLASSIFIED			1b RESTRICTIVE MARKINGS		
2a SECURITY CLASSIFICATION AUTHORITY SEP 03 1991			3 DISTRIBUTION/AVAILABILITY OF REPORT Approved for Public Release Distribution Unlimited		
2b DECLASSIFICATION/DOWNGRADING SCHEDULE					
4 PERFORMING ORGANIZATION REPORT NUMBER(S)			5 MONITORING ORGANIZATION REPORT NUMBER(S)		
6a NAME OF PERFORMING ORGANIZATION Department of Chemistry University of North Carolina		6b OFFICE SYMBOL (If applicable)	7a NAME OF MONITORING ORGANIZATION Office of Naval Research		
6c ADDRESS (City, State, and ZIP Code) Campus Box 3290 Chapel Hill, NC 27599-3290		7b ADDRESS (City, State, and ZIP Code) Department of the Navy Arlington, VA 22217			
8a NAME OF FUNDING/SPONSORING ORGANIZATION Office of Naval Research		8b OFFICE SYMBOL (If applicable)	9 PROCUREMENT INSTRUMENT IDENTIFICATION NUMBER N0014-89-J-1734		
8c ADDRESS (City, State, and ZIP Code) 800 North Quincy Street Arlington, VA 22217-5000		10 SOURCE OF FUNDING NUMBERS			
		PROGRAM ELEMENT NO	PROJECT NO	TASK NO	WORK UNIT ACCESSION NO
11 TITLE (Include Security Classification) UNCLASSIFIED: Synthesis and Characterization of Tetrathiafulvalene Charge Transfer Compounds with Copper Halides					
12 PERSONAL AUTHOR(S) Young I. Kim and William E. Hatfield					
13a TYPE OF REPORT Technical Report		13b TIME COVERED FROM TO		14. DATE OF REPORT (Year, Month, Day) 1991 August 19	
				15 PAGE COUNT 21	
16 SUPPLEMENTARY NOTATION Technical Report 44					
17 COSATI CODES			18 SUBJECT TERMS (Continue on reverse if necessary and identify by block number)		
FIELD	GROUP	SUB-GROUP	charge transfer compounds tetrathiafulvalene		
			semiconducting compounds magnetic susceptibility		
			electrical conductivities copper EPR		
19 ABSTRACT (Continue on reverse if necessary and identify by block number)					
<p>Charge transfer compounds with the general formula $(TTF)_4CuX_2$ ($TTF =$ tetrathiafulvalene; $X = Cl, Br$) may be prepared by the reaction of excess TTF with $Cu(sparteine) X_2$ in methylene chloride. The semiconducting compounds have low activation energies and relatively high electrical conductivities. Magnetic susceptibility, EPR and spectroscopic data have low activation energies and relatively high electrical conductivities. Magnetic susceptibility, EPR, and spectroscopic data provide evidence that copper has been reduced to copper(I) in the complexes, and the unpaired electrons are delocalized over columnar stacks of TTF. Reaction of TTF with $[Cu(2-aminoethylpyridine)Cl_2]$ or $[Cu(nicotine)Cl_2]$ in methylene chloride yielded $(TTF)_2CuCl_2$, while reaction of TTF with $[Cu_4OCl_6(2-picoline)_4]$ yielded $(TTF)CuCl_2$.</p>					
20 DISTRIBUTION/AVAILABILITY OF ABSTRACT <input checked="" type="checkbox"/> UNCLASSIFIED/UNLIMITED <input type="checkbox"/> SAME AS RPT <input type="checkbox"/> DTIC USERS			21 ABSTRACT SECURITY CLASSIFICATION UNCLASSIFIED		
22a NAME OF RESPONSIBLE INDIVIDUAL			22b TELEPHONE (Include Area Code)		22c OFFICE SYMBOL

OFFICE OF NAVAL RESERACH

Contract N0014-89-J-1734

R&T Code 4135007---05

TECHNICAL REPORT NO. 44

**Synthesis and Characterization of Tetrathiafulvalene
Charge Transfer Compounds with Copper Halides**

by

Young I. Kim and William E. Hatfield

The Department of Chemistry, The University of North
Carolina at Chapel Hill, North Carolina, 27599-3290, U.S.A.

Prepared for Publication in

INORGANICA CHIMICA ACTA

Reproduction in whole or in part is permitted for
any purpose of the United States Government

*This document has been approved for public release
and sale, its distribution is unlimited.

*This statement should also appear in Item 3 of Document Control
Data - DD Form 1473. Copies of the form are available from the
cognizant contract administrator.

01 8 00 000

91-09336



**Synthesis and Characterization of Tetrathiafulvalene
Charge Transfer Compounds with Copper Halides**

YOUNG INN KIM^{*} and WILLIAM E. HATFIELD^{*}

Department of Chemistry, University of North Carolina

at Chapel Hill, Chapel Hill, North Carolina 27599-3290 (U. S. A.)

Abstract

Charge transfer compounds with the general formula $(\text{TTF})_n\text{CuX}_2$ (TTF = tetrathiafulvalene; X = Cl, Br) may be prepared by the reaction of excess TTF with $\text{Cu}(\text{sparteine})\text{X}_2$ in methylene chloride. The semiconducting compounds have low activation energies and relatively high electrical conductivities. Magnetic susceptibility, EPR, and spectroscopic data provide evidence that copper has been reduced to copper(I) in the complexes, and that the unpaired electrons are delocalized over columnar stacks of TTF. Reaction of TTF with $[\text{Cu}(2\text{-aminoethylpyridine})\text{Cl}_2]$ or $[\text{Cu}(\text{nicotine})\text{Cl}_2]$ in methylene chloride yielded $(\text{TTF})_2\text{CuCl}_2$, while reaction of TTF with $[\text{Cu}_4\text{OCl}_6(2\text{-picoline})_4]$ yielded $(\text{TTF})\text{CuCl}_2$.

^{*}Current Address: Department of Chemical Education, Pusan
National University, Pusan, 609-735, KOREA

Introduction

Tetrathiafulvalene (TTF) and its analogues have been used as electron donors to form highly electroconductive charge transfer complexes [1] with the notable example being TTF-TCNQ [2] (TCNQ = tetracyanoquinodimethane). Other electroconductive charge transfer complexes with TTF contain metal halides [3], simple halides [4], and pseudohalides [5] as electron acceptors. Some transition metal complexes with typical organic ligands have been used as acceptors [6]. Since the reduction potential of transition metal complexes can be readily modified by the selection of ligands and ligand substitution, such complexes are very good candidates for use in designed syntheses. Here the goal is to prepare charge transfer complexes with fractional electron transfer, since partial electron transfer is among the factors that lead to high electrical conductivity.

Recently Inoue and coworkers [7] have reported a variety of conductive TTF-copper halide compounds by the direct reaction of excess TTF with copper halides in various solvents. They pointed out that the composition of the compounds $(\text{TTF})_n\text{CuX}_2$ ($\text{X} = \text{Cl}$ and Br) was dependent on the solvent employed in the reaction. We describe herein the reaction of TTF with some copper(II) coordination compounds, and give the results of the characterization by magnetic, electroconductive, and spectroscopic techniques of the new compounds formed from the reaction of TTF with $\text{Cu}(\text{sp})\text{Cl}_2$ [$\text{sp} = (-), \text{-sparteine}$] [8].

For	
GI	<input checked="" type="checkbox"/>
ed	<input type="checkbox"/>
tion	<input type="checkbox"/>
on/	
ity Codes	
and/or	
pecial	



A-1

Experimental

Synthesis

A solution of 0.2 g of TTF in 10 ml of methylene chloride was added to a methylene chloride (10 mL) solution of 0.5 g of $\text{Cu}(\text{sp})\text{Cl}_2$. TTF was obtained from Stream Chemicals and was used without further purification. $\text{Cu}(\text{sp})\text{Cl}_2$ was prepared by a method that has been described previously [9]. The mixture of TTF and $\text{Cu}(\text{sp})\text{Cl}_2$ was stirred for 30 minutes at room temperature. The deep purple solid product that precipitated was filtered off and washed with methylene chloride. The product was dried in vacuum at room temperature. The purple bromide complex was synthesized by a similar method. Element analysis were obtained from Galbraith Laboratories, Inc., in Knoxville, Tennessee.

$(\text{TTF})_4\text{CuCl}_2$ Calcd: C, 30.28; H, 1.69. Found: C, 30.89; H, 2.02.

$(\text{TTF})_4\text{CuBr}_2$ Calcd: C, 27.69; H, 1.55. Found: C, 27.28; H, 1.63.

Reaction of TTF with $[\text{Cu}(2\text{-aminoethylpyridine})\text{Cl}_2]$ or $[\text{Cu}(\text{nicotine})\text{Cl}_2]$ in methylene chloride yielded $(\text{TTF})_2\text{CuCl}_2$, while the reaction of TTF with $[\text{Cu}_4\text{OCl}_6(2\text{-picoline})_4]$ yielded $(\text{TTF})\text{CuCl}_2$. These known compounds were characterized by chemical analysis and physical measurements. Our results were consistent with known properties, and they will not be discussed here.

Physical measurements

Four probe d.c. electrical resistivities were measured on pressed pellets, 1.3 cm in diameter and about 0.1 cm in thickness, using a CTI-Cryogenics 21SC Cryodine Cryocooler and the Van der Pauw technique [10]. The temperature was controlled

by a Lake Shore Cryotronics DRS 80C temperature controller. EPR spectral measurements were made on powdered samples at 77 K and at room temperature by using a Varian E-109 X-band spectrometer. The field strength was calibrated using DPPH ($g = 2.0037$). Magnetic susceptibility data were collected from 77 K to room temperature by using the Faraday method with a Cahn 2000 electrobalance. The magnetometer was calibrated with $\text{HgCo}(\text{NCS})_4$ [11]. Infrared spectra ($400\text{--}1500\text{ cm}^{-1}$) were obtained using Nujol mulls on cesium iodide plates with a Nicolet Model 200X FT-IR spectrophotometer. Electronic spectra in the uv-visible range were obtained with a Hewlett-Packard 8451A spectrophotometer on Nujol mulls mounted between quartz plates. Cyclicvoltammograms were recorded on a Princeton Applied Research (PAR) model 173 Cyclic Voltammograph at $\text{pH} = 7.01$ in aqueous solutions at a scanning rate of 200 mV/sec . All potentials are reported versus the saturated calomel electrode (SCE).

Results and Discussion

$(\text{TTF})_4\text{CuX}_2$ (where $\text{X}=\text{Cl}$ and Br) were obtained by the reaction of TTF and $\text{Cu}(\text{sp})\text{X}_2$ in methylene chloride. The composition of the products was always four TTF per copper, and the composition was independent of the amount of TTF as long as an excess was used in methylene chloride.

Electrical Properties. The electrical conductivity of $(\text{TTF})_4\text{CuX}_2$ was measured by using the van der Pauw four-probe d. c. method in the range of 70 K to 300 K. As shown in Table I,

the powdered samples of each compound exhibits high electrical conductivities at room temperature. The conductivities are higher than those of halide salts of TTF [4a,12] which are known to have columnar structures of partially oxidized TTF. The resistivities of both compounds increases with decreasing temperature in the range of 70 K to 300 K. The negative coefficient $d\rho/dT < 0$ reflects the intrinsic semiconductor behavior given by $\rho = \rho_0 \exp(E_g/kT)$. Similar behavior has been found in TTF salts with several transition metal complex anions [3a,6,7]. A plot of $\log \rho$ vs $1/T$ for $(\text{TTF})_4\text{CuBr}_2$ showed a small inflection near 230 K. the data above 230 K may be fit by the above equation with $E_g = 0.21 \times 10^{-1}$ eV, and the data below 230 K may be fit with $E_g = 2.40 \times 10^{-1}$ eV. The data for $(\text{TTF})_4\text{CuCl}_2$ did not exhibit an inflection.

The mobility model for electrical resistivity is given by [13]

$$\rho(T) = A \cdot T^\alpha \exp(E_g/kT)$$

The Boltzmann term accounts for an activated generation of charge carriers in a narrow-band gap semiconductor. The preexponential term is associated with a temperature dependence of the mobility. The data for $(\text{TTF})_4\text{CuX}_2$ was fitted by the above equation. The best-fit parameters are $A = 8.58 \times 10^1$, $\alpha = 1.48$, and $E_g = 1.88 \times 10^{-1}$ eV for $(\text{TTF})_4\text{CuCl}_2$, and $A = 1.12 \times 10^2$, $\alpha = 1.47$, and $E_g = 1.28 \times 10^{-1}$ eV for $(\text{TTF})_4\text{CuBr}_2$. The values for the parameter α are nearly equal to the theoretical value ($\alpha = 1.5$) for the temperature dependence of mobility for the small polaron

mechanism [14]. The mobility model has been used to describe the conductivity in other electrical conductors including a series of porphyrinic molecular metals [15-17].

Magnetic Properties. EPR spectra for powdered samples were obtained both at room temperature and at 77 K. At room temperature $(\text{TTF})_4\text{CuCl}_2$ exhibits a slightly unsymmetrical spectral band with the anisotropic g values, $g_{\parallel} = 2.017$ and $g_{\perp} = 2.010$, while $(\text{TTF})_4\text{CuBr}_2$ exhibits a symmetric EPR spectral band with $\langle g \rangle = 2.007$. The EPR line shapes for both compounds exhibit good resolution of parallel (g_{\parallel}) and perpendicular (g_{\perp}) components at 77 K.

The average $\langle g \rangle$ values of $(\text{TTF})_4\text{CuCl}_2$ and $(\text{TTF})_4\text{CuBr}_2$ at 77 K are 2.010 and 2.003, respectively, values which are close to the g value of the TTF^+ ion in solution ($g = 2.00838$) [18]. These values are also comparable to those of TTF-halides and TTF-pseudohalides as shown in Table II. The similarity of the g values indicate that the odd electrons reside on TTF in $(\text{TTF})_4\text{CuX}_2$. A signal attributable to Cu(II) was not detected in any case, an indication that the copper ions in $(\text{TTF})_4\text{CuX}_2$ are in the diamagnetic Cu(I) state. We can not rule out the possibility that some electron density is delocalized over several copper ions resulting in a broad undetectable band, but we think this is unlikely.

The magnetic susceptibility data also reflect the diamagnetism of copper in each compound. The temperature dependence of the magnetic susceptibilities from 80 to 300 K are shown in Figure 1. The magnetic susceptibility of $(\text{TTF})_4\text{CuX}_2$ increases somewhat as the temperature decreases, but the data are not well described by the Curie law $\chi(T) = C/T$. Weak

paramagnetism is well known in highly conducting molecular metals [19]. Here electron delocalization along TTF columns in $(\text{TTF})_4\text{CuX}_2$ requires the use of the expression $\chi(T) = C/T^a$, where a is less than 1. The magnetic susceptibility data may be described by the power law with a equal to 0.46 for $(\text{TTF})_4\text{CuCl}_2$ and 0.37 for $(\text{TTF})_4\text{CuBr}_2$. The best-fit C values are 9.93×10^{-3} and 3.73×10^{-3} for $(\text{TTF})_4\text{CuCl}_2$ and $(\text{TTF})_4\text{CuBr}_2$, respectively. A power law has been used to describe the magnetic properties of such compounds as quinolinium-TCNQ [20] and the tetramethyl-p-phenylene diamine salt (TMPD)-TCNQF₄ [21].

The presence of TTF columns in $(\text{TTF})_4\text{CuX}_2$ is supported by the narrow peak-to-peak line widths (ΔH_{pp}) in the EPR spectra. The line width of the EPR signals at room temperature are 17 gauss for $(\text{TTF})_4\text{CuCl}_2$ and 12 gauss for $(\text{TTF})_4\text{CuBr}_2$. This is somewhat larger than that of TTF•TCNQ (~6G) [22] and similar to that of TTF•SeCN (~15G) [23]. Line widths of five to fifteen gauss are attributed to spin-orbit interaction of sulfur in TTF columnar chains [12]. The extremely broad line width, ~180-200 gauss for TTF.I_{0.7} has been explained as arising from back charge transfer of electrons between TTF and iodide [12].

Spectroscopic Properties. The infrared spectra of $(\text{TTF})_4\text{CuX}_2$ consist of very broad bands extending from 1000 to 4000 cm^{-1} . These broad absorptions, which arise from the band structure of these semiconductors [24], mask many of the vibrational bands. Vibrational bands of TTF were assigned to the absorptions at 827 cm^{-1} (ν_{16}), 1280 cm^{-1} (ν_{23}) for $(\text{TTF})_4\text{CuCl}_2$ and

828 cm^{-1} (ν_{16}), 816 cm^{-1} (ν_{25}), and 1242 cm^{-1} (ν_{23}) for $(\text{TTF})_4\text{CuBr}_2$ by comparison with reported spectra [25]. The ν_{16} vibrational mode is associated with the stretching of the C-S bond in the five membered ring of the TTF molecule. The ν_{23} band arises from CCH, and the ν_{25} band arises from the ring SCC bend. The C-S band as well as the C=C modes are expected to be shifted as a result of variation of bond orders and bond lengths due to the oxidation of TTF. The observed values of ν_{16} in $(\text{TTF})_4\text{CuX}_2$ are higher than the value of 781 cm^{-1} for the TTF molecule and lower than 836 cm^{-1} for TTF in $\text{TTF}\cdot\text{Br}$ [25]. Inoue et al. [7] pointed out a linear relationship between the shift in the ν_{16} band and the charge on TTF in TTF-copper halides compounds. Since ν_{16} is shifted here, the infrared results confirm partial ionization of TTF in $(\text{TTF})_4\text{CuX}_2$.

Ultraviolet-visible spectra of solid samples mulled in Nujol revealed absorption maxima (λ_{max}) at 246 and 362 nm for $(\text{TTF})_4\text{CuCl}_2$ and at 270, 364, and 486 nm for $(\text{TTF})_4\text{CuBr}_2$. The low energy absorption band above 500 nm which is often found in conducting TTF compounds [26] was part of the broad background and a specific wavelength could not be assigned.

Electrochemistry. Cyclic voltammograms of $(\text{TTF})_4\text{CuX}_2$ and $\text{TTF}\cdot\text{I}_{0.7}$ [27] were recorded in aqueous solution at pH = 7.01 vs a saturated calomel electrode (SCE). The results are summarized in Table I. $(\text{TTF})_4\text{CuCl}_2$ exhibited three peaks; at +0.13 V for the $\text{Cu}^{2+}/\text{Cu}^+$ couple, at +0.43 V for the TTF^+/TTF couple, and at +0.79 V for the $\text{TTF}^{2+}/\text{TTF}^+$ couple. Similar results were observed for $(\text{TTF})_4\text{CuBr}_2$. The potentials ($E_{1/2}$) reported for the couples were

estimated by averaging the anodic and cathodic peak potentials. The $E_{1/2}$ potential for the $\text{Cu}^{2+}/\text{Cu}^+$ couple is nearly equal to the standard reduction potential, $\text{Cu(II)} \rightarrow \text{Cu(I)}$ (+0.159 V vs SCE) in aqueous solution [28]. Two peaks were found in $\text{TTF} \cdot \text{I}_{0.7}$, which are the redox potentials for the TTF couples. The experimentally observed $E_{1/2}$ value of $\text{TTF} \cdot \text{I}_{0.7}$ were +0.42 and +0.73 V. These were assigned to the TTF^+/TTF and $\text{TTF}^{2+}/\text{TTF}^+$ couples, respectively and support the assignment in the case of $(\text{TTF})_4\text{CuCl}_2$. TTF in CH_3CN solution also exhibits two reversible redox waves at +0.33 (TTF^+/TTF) and at 0.70 V ($\text{TTF}^{2+}/\text{TTF}^+$) vs SCE [29]. Thin films of TTF polymer also show two waves at nearly the same potentials [30]. The cyclic voltammograms were scanned several times and there was no change in the potentials. This is good evidence that the couples are reversible.

Correlation of electrical conductivity with redox potentials has been discussed [31,32]. With the assumption that electron transfer calculated from solution redox potentials parallels that in the solid phase, it may be concluded that charge transfer compounds with low resistivities ($\rho < 0.1 \text{ ohm-cm}$) will result from the combination of moderately strong acceptors with moderately strong donors. Based on studies of TTF complexes with substituted TCNQ, Wheland [32] proposed that the equilibrium constant is related to the difference in redox potentials by relationship $\log K = E_{1A} - E_{1D}/0.059$, where E_{1A} is the redox potential of acceptor and E_{1D} is that of donor. He noted that electron acceptors which had low K values ($K = 10^{0.2}$ to 10^{-4}) gave highly conducting complexes, whereas acceptors with larger K values ($K = 10^{0.2}$ to 10^{+6}) gave poor conductors. Small K values

indicate incomplete electron transfer, and large K values indicate complete electron transfer. A value of $K = \sim 10^{-5}$ was calculated by using the above equation for $(\text{TTF})_4\text{CuCl}_2$ and $(\text{TTF})_4\text{CuBr}_2$. In these compounds there is complete reduction of copper and one electron per four TTF molecules.

Conclusions. Charge transfer compounds $(\text{TTF})_n\text{CuX}_2$ prepared from TTF and $\text{Cu}(\text{Sp})\text{X}_2$ exhibit semiconductor behavior. Magnetic and spectroscopic data reveal that copper(II) has been reduced to copper(I) and that the odd electron is delocalized on TTF molecules in columnar stacks. Orbital overlap between TTF molecules give rise to band formation, low activation energies, and relatively high electrical conductivities. Furthermore, reaction of TTF with copper(II) complexes may provide a convenient method for the preparation of $(\text{TTF})_n\text{CuX}_2$ charge transfer compounds with varying ratios of TTF to copper halide.

Acknowledgement

This research was partially funded by the Office of Naval Research.

References

- (1) See, for example, W. E. Hatfield, Molecular Metals, Plenum Press, New York, 1979.
- (2) (a) J. Ferraris, D. O. Cowan, V. V. Walatka, Jr., and J. H. Perlstein, J. Am. Chem. Soc., **95** (1973) 948. (b) L. B. Coleman, M. J. Cohen, D. J. Sandman, F. G. Yamagishi,

- A. F. Garito, and A. J. Heeger, Solid State Commun. **12** (1973) 1125.
- (3) (a) K. Kondo, G. Matsubayashi, T. Tanaka, H. Yoshioka, and K. Nakatsu, J. Chem. Soc., Dalton Trans. (1984) 379.
(b) T. J. Kistenmacher, M. Rossi, C. C. Chiang, R. P. Van Duyne, and A. R. Siedle, J. Am. Chem. Soc. **102** (1980) 3604. (c) E. I. Zhilyaeva, R. Lyubovskaya, M. L. Khidekel, M. S. Ioffe, and T. M. Moravskaya, Transition Metal Chem. **5** (1980) 189.
- (4) (a) F. Wudl, D. Wobschall, and E. J. Hubnagel, J. Am. Chem. Soc., **94** (1972) 670. (b) B. A. Scott, S. J. Laplaca, J. B. Torrance, B. D. Silverman, and B. Welber, B. J. Am. Chem. Soc., **99** (1977) 6631.
- (5) F. Wudl, D. E. Schafer, W. M. Walsh, Jr., L. W. Rupp, F. J. Disalvo, J. V. Wasycyak, M. L. Kaplan, and G. A. Thomas, J. Chem. Phys., **66** (1977) 277.
- (6) K. Ueyama, A. Tanaka, G. Matsubayashi, and T. Tanaka, Inorg. Chimica Acta., **97** (1985) 201.
- (7) (a) M. Inoue, and M. B. Inoue, J. Chem. Soc., Chem. Commun., (1985) 1043. (b) M. Inoue, M. Inoue, Q. Fernando, and K. W. Nebesny, Inorg. Chem., **25** (1986) 3976. (c) M. B. Inoue, C. Cruy-Vayquey, M. Inoue, Q. Fernando, and K. W. Nebesny, Inorg. Chem., **26** (1987) 1462.
- (8) S. N. Choi, R. D. Bereman, and J. R. Wasson, Inorg. Nucl. Chem., **37** (1975) 2087.
- (9) S. F. Mason and R. D. Peacock, J. Chem. Soc., Dalton Trans.

- (1973) 226.
- (10) J. L. Van der Pauw, Philips Tech. Rev., 20 (1959) 220.
- (11) D. B. Brown, V. H. Crawford, J. W. Hall, and W. E. Hatfield, J. Phys. Chem., 81 (1977) 1303.
- (12) R. B. Somoano, A. Gupta, V. Hadek, T. Datta, M. Jones, R. Deck, and A. M. Hermann, J. Chem. Phys., 63 (1975) 4970.
- (13) A. J. Epstein, E. M. Conwell, D. J. Sandman, and J. S. Miller, Solid State Commun., 23 (1977) 355.
- (14) P. F. Weller, Solid State Chemistry and Physics, Vol. 1, Marcel Dekker Inc., 1973.
- (15) T. E. Phillips, R. P. Scaringe, B. M. Hoffman, and J. A. Ibers, J. Am. Chem. Soc., 102 (1980) 3435.
- (16) J. Martinsen, L. J. Pace, T. E. Phillips, B. M. Hoffman, and J. A. Ibers, J. Am. Chem. Soc., 104 (1982) 83.
- (17) J. Padilla, J. Ph. D. Dissertation, University of North Carolina at Chapel Hill, 1988.
- (18) F. Wudl, G. M. Smith, and E. J. Hufnagel, E. J. J. Chem. Soc., Chem. Commun., (1970) 1453.
- (19) B. M. Hoffman and J. A. Ibers, Acc. Chem. Res., 16 (1983) 15.
- (20) L. C. Tippie and W. G. Clark, Phy. Rev., B23 (1981) 5846.
- (21) W. E. Hatfield and L. W. ter Haar, Ann. Rev. Mat. Sci., 12 (1982) 177.
- (22) Y. Tomkiewicz, B. A. Scott, L. J. Tao, and R. S. Title, Phy. Rev. Let., 32 (1974) 1363.

- (23) Y. Tomkiewicz and E. M. Engler, Bull. Am. Phys. Soc.,
20 (1975) 479.
- (24) M. J. Rice, L. Pietronero, and P. Bruesch, P. Solid State
Commun., 21 (1977) 757.
- (25) R. Boyio, I. Zanon, A. Girlando, and C. Pecile, J. Chem.
Phys., 71 (1979) 2282.
- (26) J. B. Torrance, B. A. Scott, B. Welber, F. B. Kaufman,
and P. E. Seiden, Phy. Rev., B19 (1979) 730.
- (27) L. R. Melby, Can. J. Chem., 43 (1965) 1448.
- (28) A. J. Bard, R. Parsons, and J. Jordans, J. Standard
Potentials in Aqueous Solution, Marcel Dekker, Inc., New
York and Basel, 1985.
- (29) D. L. Coffen, J. Q. Chambers, D. R. Williams, P. E.
Garrett, and N. D. Canfield, J. Am. Chem. Soc., 93 (1971)
2258.
- (30) F. B. Kaufman, A. H. Schroeder, E. M. Engler, S. R. Kramer,
and J. Q. Chambers, J. Am. Chem. Soc., 102 (1980) 483.
- (31) J. B. Torrance, Acc. Chem. Res., 12 (1979) 79.
- (32) R. C. Wheland, J. Am. Chem. Soc., 98 (1976) 3926.

TABLE I. Physical Properties of TTF-CuX₂ Complexes

Compounds	Electrical		Magnetic		Redox Potential	
	Conductivity (σ) at room temperature (S cm ⁻¹)		Susceptibility (χ) at room temperature (emu mol ⁻¹)		vs SCE at pH = 7.01 (volt)	
(TTF) ₄ CuCl ₂	1.69 x 10		7.44 x 10 ⁻⁴ (1.28) *			+0.13
						+0.43
						+0.79
(TTF) ₄ CuBr ₂	6.78 x 10		4.79 x 10 ⁻⁴ (1.08)			+0.13
						+0.44
						+0.78

*Magnetic moments (B.M.) at room temperature are listed in parenthesis.

TABLE 2. Anisotropic g values and linewidths of EPR spectra of some TTF complexes.

Compound	Anisotropic g values		EPR line width Gauss	Remarks
	g_{\parallel}	g_{\perp}		
TTF•ClX	2.0033	2.0097	9	(1)
TTF•Br _{0.7}	2.0025	2.0100	40-52	(2)
TTF•I _{0.7}	2.0050	2.0128	180-200	(2)
(TTF) ₁₁ (SCN) ₆	2.0020	2.0088	11	(3)
(TTF) ₁₁ (SeCN) ₆	2.0022	2.0086	15	(3)
(TTF) ₄ CuCl ₂	1.9989	2.0049	17	this work
(TTF) ₄ CuBr ₂	2.0002	2.0074	12	this work

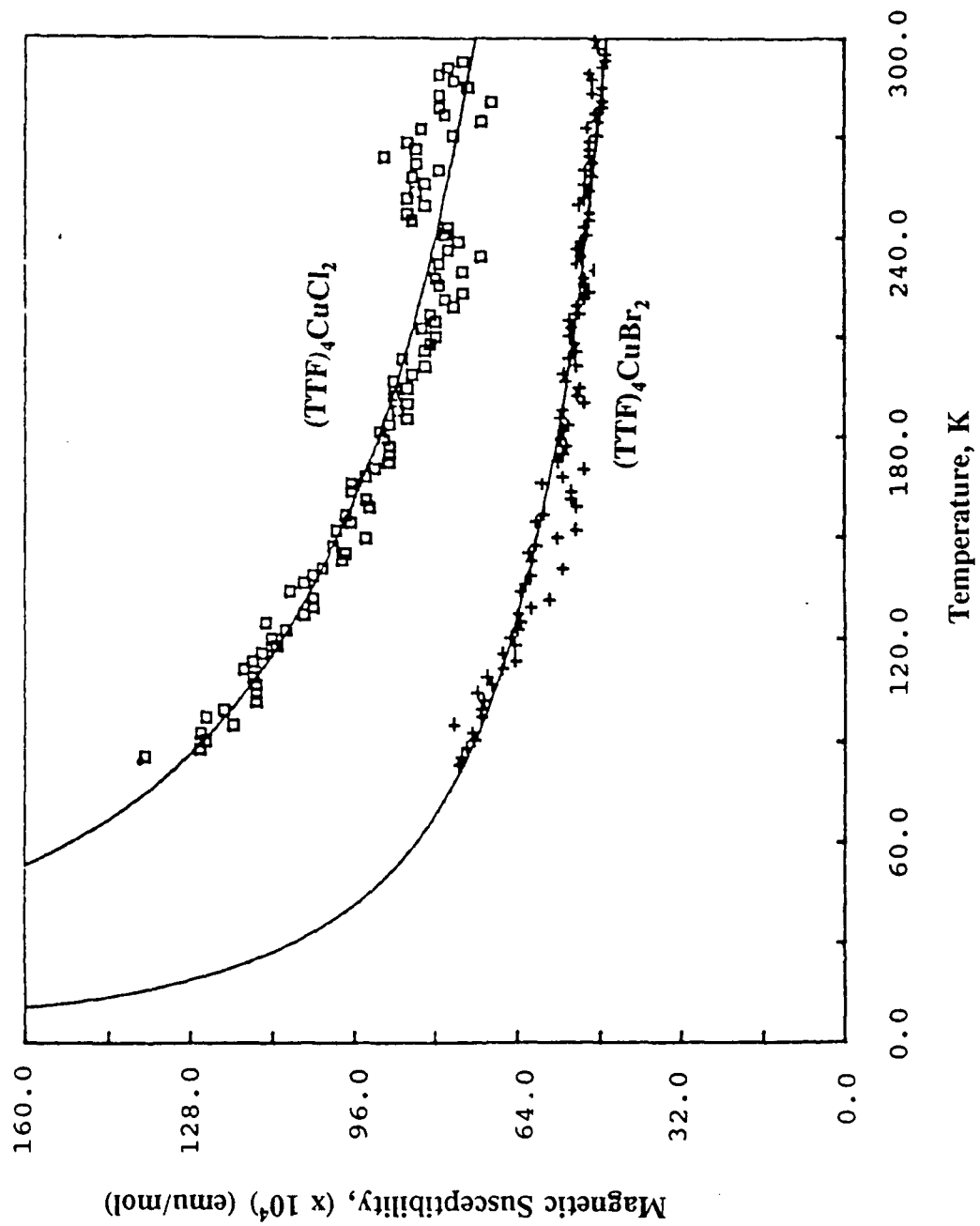
(1) Ref 4(b)

(2) Ref. 5 and T. Sugano and H. Kuroda, Chem. Phys. Lett., **47** (1977) 92.

(3) Ref. 5

FIGURE CAPTION

Figure 1. Temperature dependence of magnetic susceptibility of $(\text{TTF})_4\text{CuCl}_2$ and $(\text{TTF})_4\text{CuBr}_2$.



TECHNICAL REPORT DISTRIBUTION LIST, GENERAL

	<u>No. Copies</u>		<u>No. Copies</u>
Office of Naval Research Chemistry Division, Code 1113 800 North Quincy Street Arlington, VA 22217-5000	3	Dr. Ronald L. Atkins Chemistry Division (Code 385) Naval Weapons Center China Lake, CA 93555-6001	1
Commanding Officer Naval Weapons Support Center Attn: Dr. Bernard E. Doua Crane, IN 47522-5050	1	Chief of Naval Research Special Assistant for Marine Corps Matters Code OOMC 800 North Quincy Street Arlington, VA 22217-5000	1
Dr. Richard W. Drisko Naval Civil Engineering Laboratory Code L52 Port Hueneme, California 93043	1	Dr. Bernadette Eichinger Naval Ship Systems Engineering Station Code 053 Philadelphia Naval Base Philadelphia, PA 19112	1
Defense Technical Information Center Building 5, Cameron Station Alexandria, Virginia 22314	2 <u>high</u> <u>quality</u>	Dr. Sachio Yamamoto Naval Ocean Systems Center Code 52 San Diego, CA 92152-5000	1
David Taylor Research Center Dr. Eugene C. Fischer Annapolis, MD 21402-5067	1	David Taylor Research Center Dr. Harold H. Singerman Annapolis, MD 21402-5067 ATTN: Code 283	1
Dr. James S. Murday Chemistry Division, Code 6100 Naval Research Laboratory Washington, D.C. 20375-5000	1		

ORGANOELEMENT CHEMISTRY - Distribution List

Professor O. T. Beachley, Jr.
Department of Chemistry
State University of New York
Buffalo, NY 14214
R&T Code 4135002

Professor Herbert C. Brown
Purdue University
Department of Chemistry
West Lafayette, IN 47907
R&T Code 4135011

Professor Steven L. Buchwald
Department of Chemistry
Massachusetts Institute of Technology
Cambridge, MA 02139
R&T Code 4135014

Professor William E. Hatfield
Department of Chemistry
University of North Carolina
Chapel Hill, NC 27514
R&T Code 4135007

Professor M. Frederick Hawthorne
Department of Chemistry
405 Hilgard Avenue
University of California
Los Angeles, CA 90024
R&T Code 4135004

Professor Robert H. Neilson
Department of Chemistry
Texas Christian University
Fort Worth, TX 76843
R&T Code 4135005

Professor Kurt Niedenzu
Department of Chemistry
University of Kentucky
Lexington, KY 40506
R&T Code 4135003

Professor Richard L. Wells
Department of Chemistry
Duke University
Durham, NC 27706
R&T Code 4135008